



# Engineering MoS<sub>2</sub> Nanosheets Anchored on Metal Organic Frameworks Derived Carbon Polyhedra for Superior Lithium and Potassium Storage

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### Specialty section:

This article was submitted to  
Electrochemical Energy Conversion  
and Storage,  
a section of the journal  
Frontiers in Energy Research

**Received:** 16 October 2019

**Accepted:** 18 November 2019

**Published:** 03 December 2019

### Citation:

Rui B, Li J, Chang L, Wang H, Lin L,  
Guo Y and Nie P (2019) Engineering  
MoS<sub>2</sub> Nanosheets Anchored on Metal  
Organic Frameworks Derived Carbon  
Polyhedra for Superior Lithium and  
Potassium Storage.  
Front. Energy Res. 7:142.  
doi: 10.3389/fenrg.2019.00142

Metal organic frameworks (MOFs) have exhibited promising potential as a new platform for the preparation of porous functional materials for energy storage application. Herein, we report an approach to synthesize ultrathin MoS<sub>2</sub> nanosheets decorated cobalt nanoparticles-containing porous carbon polyhedral with dense nitrogen doped carbon nanotubes (CNTs) backbone, where the ZIF-67 is used as the carbon and nitrogen sources for the growth of CNTs and the self-template for the carbon polyhedron. The composite shows a unique structure in which ultrathin MoS<sub>2</sub> nanosheets are uniformly coated on the ZIF-67 derived carbon polyhedron. Owing to the synergistic effect from composition, morphology, and robust hollow structure composed of interconnected nitrogen doped CNTs networks and cobalt particles, the resulting ZIF-67-C@MoS<sub>2</sub> nanocomposites display high specific capacity of 568.5 mAh g<sup>-1</sup> at a current density of 0.2 A g<sup>-1</sup>, superior rate capability (291.9 mAh g<sup>-1</sup> at 5 A g<sup>-1</sup>), and excellent cycling stability for over 500 cycles as lithium ion batteries anodes. The material also exhibits superior performance as new anodes for potassium ion batteries. The results provide a facile strategy for large-scale synthesis of high performance anode materials through low cost perspective and new insights for designing MOF-derived functional materials with unique structures for energy storage.

**Keywords:** metal-organic frameworks, MoS<sub>2</sub>, carbon polyhedra, lithium ion batteries, potassium ion batteries

## INTRODUCTION

The need for large-scale energy storage on the smart grid and electric vehicles is growing rapidly. Electrochemical energy storage systems are highly desirable because they can offer high energy density, rate capability and rapid response (Dunn et al., 2011; Choi and Aurbach, 2016; Le et al., 2017; Liu et al., 2017, 2018; Peng et al., 2017; Sun et al., 2018; Davies et al., 2019; Ding et al., 2019). In the last two decades, lithium ion batteries (LIBs) have become increasingly popular because they possess superior energy density, light weight and long cycle life (Kovalenko et al., 2011; Wang et al., 2013; Li et al., 2016; Nie et al., 2017a, 2018; Griffith et al., 2018; Sun et al., 2019). The ever-growing demands for large scale applications of LIBs have opened a new area in advanced materials research.

It is well-established that metal organic frameworks has gained enormous interests. As a new type of porous crystal material, it is composed of organic ligands and inorganic metal ions. Because of the high specific surface area, high porosity, and adjustable pore size, it widely used in many fields (Yunhong et al., 2017; Chen et al., 2018a; Mínguez Espallargas and Coronado, 2018) including catalysis (Chen et al., 2018c; Zhang et al., 2018c), hydrogen storage (Basdogan and Keskin, 2015; Kim et al., 2015), adsorption (Xue et al., 2013; Oveisi et al., 2018) and energy storage (Nie et al., 2014, 2015, 2017b; Wu and Lou, 2017; Xu et al., 2017; Wang et al., 2018). However, the conductivity of most MOFs is relatively low, and exhibiting low specific capacity as anode materials for lithium ion batteries (Chen et al., 2015). MOFs derived carbon materials could inherit the pristine porous structure of MOFs, and have tunable doping by heteroatoms and improved conductivity, thus providing a new platform for the synthesis of carbon based nanocomposites (Wang et al., 2017; Gadipelli et al., 2019). Jagadeesh et al. fabricated a graphitic carbon-coated cobalt for amination catalysts through pyrolysis of Co-MOFs precursors. The well-developed Co catalyst enables the effective synthesis of amines, amino acid derivatives, and more complex drug targets (Chen and Xu, 2017; Jagadeesh et al., 2017). Three-dimensional (3D) carbon nanomaterials have attracted considerable attention as promising electrodes for electrochemical energy storage application due to their unique structural characteristics of interlinked architecture (Chen et al., 2017a). 3D porous MnO/C-N nanocomposites have been prepared by using renewable rapeseed pollen via a facile immersion-annealing route (Chen et al., 2017c), which showed excellent electrochemical performance in terms of high specific capacity, long cycling life over 400 cycles and super rate capability. Chen et al. reported the synthesis of hollow particle-based nitrogen-doped carbon nanofibers by embedding ultrafine zeolitic imidazolate framework (ZIF-8) nanoparticles into electrospun polyacrylonitrile fibers followed by carbonization (Chen et al., 2017b). When used as electrodes for supercapacitors, the hierarchical porous nanofibers exhibits superior electrochemical performance with remarkable specific capacitances, high energy/power density as well as outstanding cycling stability for 10,000 cycles at 5 A g<sup>-1</sup>.

Since the successful preparation of single layer graphene in 2004 (Novoselov et al., 2004), two-dimensional (2D) materials including transition metal dichalcogenides (TMDs), metal coordination polymer, layered transition metal oxide, have been extensively explored for energy storage application owing to outstanding mechanical flexibility and planar geometry (Zhang et al., 2018b). Among 2D transition metal dichalcogenides, MoS<sub>2</sub> facilitates the insertion and extraction of alkali ions because of its weak inter-layer van der Waals force and a high theoretical specific capacity due to its multi-electron transfer mechanism (Tian et al., 2019). In addition, the material has the merit of low cost, thus has received considerable interests as promising anode materials for alkali-ion batteries (Jiao et al., 2017; Xie et al., 2017; Xu et al., 2018; Jiang et al., 2019). However, MoS<sub>2</sub> suffers from the issues of fast capacity decay caused by mechanical strain induced during charge/discharge process and low rate capability because of intrinsically poor conductivity. When used as a host

for lithium storage, besides intercalation, MoS<sub>2</sub> can be converted into Mo particles and Li<sub>2</sub>S with additional substantial capacity and a 103% volume expansion (Xiao et al., 2010; Stephenson et al., 2014; Zhu et al., 2014). Li intercalation changes the Mo coordination from trigonal-prismatic to octahedral coordination. Furthermore, the poor electronic/ionic conductivity between two adjacent S–Mo–S layers also limits their applications.

Inspired by the above considerations, we herein report a strategy to synthesize ultrathin MoS<sub>2</sub> decorated cobalt nanoparticles-containing porous carbon polyhedral with dense carbon nanotubes backbone, where the ZIF-67 is used as the carbon and nitrogen sources for the growth of carbon nanotubes and the self-template for the polyhedron structure. The material shows a unique structure in which ultrathin MoS<sub>2</sub> nanosheets are uniformly coated on the ZIF-67 derived carbon polyhedron. With many advantageous features in composition, morphology, and robust hollow structure with nitrogen doped CNTs, the resulting ZIF-67-C@MoS<sub>2</sub> nanocomposite displays high specific capacity of 568.5 mAh g<sup>-1</sup> at 0.2 A g<sup>-1</sup>, superior rate capability and excellent cycling stability (350.4 mAh g<sup>-1</sup> upon 500 cycles) as lithium ion batteries anodes. It also exhibited superior performance as new anodes for potassium ion batteries.

## EXPERIMENTAL

### Materials Synthesis

#### Synthesis of ZIF-67-C

ZIF-67-derived carbon materials were synthesized as follows. Firstly, 2-methylimidazole (1.97 g) and Co(NO<sub>3</sub>)<sub>2</sub>·6H<sub>2</sub>O (1.746 g) was dissolved in a mixed solution of 20 mL methanol and 20 mL ethanol, respectively. The above two solutions were then mixed together under continuous stirring, and the final solution was aged at room temperature for 24 h. After that, the purple precipitate was collected by centrifugation, washed in ethanol several times and dried at 60°C. For the synthesis of the ZIF-67 derived carbon, the ZIF-67 particles were heated to 350°C for 1.5 h, then 800°C for 3.5 h with a ramp rate of 2°C min<sup>-1</sup> under Ar/H<sub>2</sub> atmosphere (9:1 in vol. ratio). After cooling down to room temperature, the carbonized black powders were immersed in 0.5 M H<sub>2</sub>SO<sub>4</sub> solution for 6 h. The resulting ZIF-67 derived carbon was collected by centrifugation, washed with DI water, then dried at 60°C, the resulting carbon was denoted as ZIF-67-C.

#### Synthesis of ZIF-67-C@MoS<sub>2</sub>

The as-prepared ZIF-67 (30 mg) was added into 25 mL C<sub>6</sub>H<sub>12</sub>O<sub>6</sub> solution (0.248 g) and ultrasonication for several minutes. Then, 0.3 g Na<sub>2</sub>MoO<sub>4</sub>·2H<sub>2</sub>O and 0.6 g CH<sub>4</sub>N<sub>2</sub>S were in order added into above solution. After stirring for 1 h, the dispersion was transferred into a 50 mL Teflon-lined autoclave and heated at 200°C for 24 h. The as prepared black powder was washed with deionized water and then dried at vacuum overnight. Finally, the samples were annealed at 500°C for 4 h in an argon flow with a heating rate of 1°C min<sup>-1</sup>. The obtained sample was marked as ZIF-67-C@MoS<sub>2</sub>.

## Characterization

Powder X-ray diffraction (XRD) (Rigaku d/max PC2500) with Cu K $\alpha$  radiation was used to determine the crystal structure of the products. The morphology and microstructure of the samples were observed using field-emission scanning electron microscopy (FESEM, JSM-7800F).

## Electrochemical Measurements

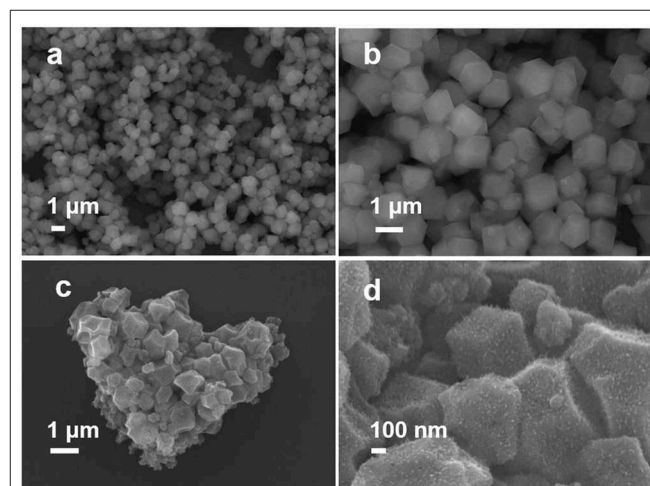
The electrochemical performances were measured by using CR2032 type coin cell. The working electrodes were prepared by slurry coating procedure. The working electrode was prepared by mixing active material, acetylene black and carboxymethyl cellulose in a weight ratio of 80:10:10 in water. Then the slurry was spread onto a copper foil and dried under vacuum at 70°C for 12 h. The cell assembly was performed in an argon-filled glove box. For lithium ion batteries, the electrolyte used was 1 M LiPF<sub>6</sub> that was dissolved in a mixture of ethylene carbonate (EC) and dimethyl carbonate (DMC) (1:1 v/v) with 5% fluoroethylene carbonate (FEC). Pure lithium metal was used as counter electrode and a polypropylene membrane as separator. For potassium ion batteries, the electrolyte was 0.8 M KPF<sub>6</sub> dissolved in a mixed solution of ethylene carbonate (EC) and diethyl carbonate (DEC) (1:1, vol%). Potassium metal was used as counter electrode, glass microfiber membrane as separators, respectively. Galvanostatic measurements were performed under different current densities between 0.01 and 3.0 V using a CT2001A cell test instrument (LAND Electronic Co.). Cyclic voltammetry (CV) was measured using an electrochemical workstation in the voltage range of 0.01–3.0 V.

## RESULTS AND DISCUSSION

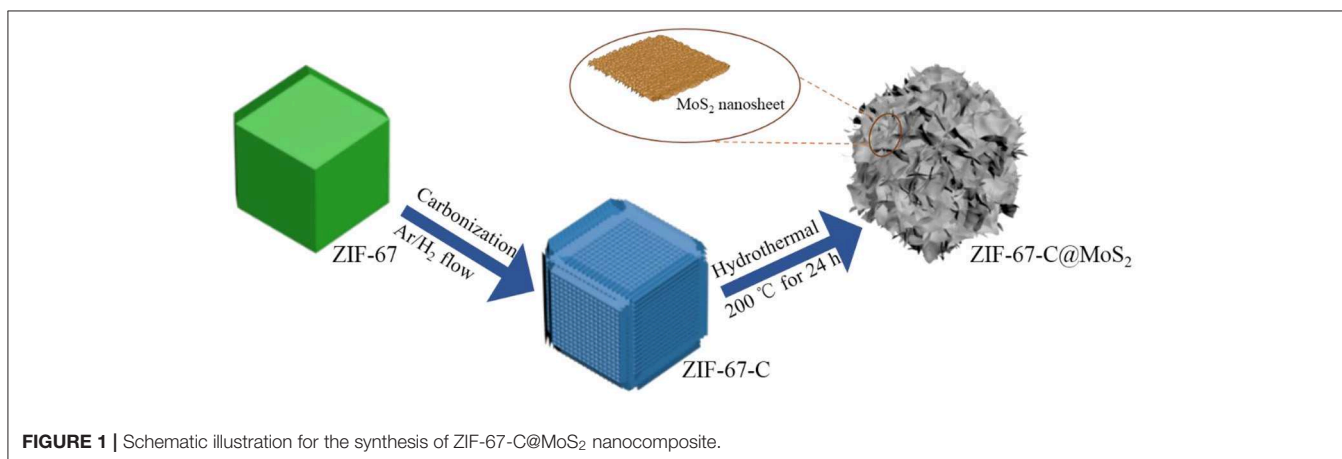
We designed and synthesized ultrathin MoS<sub>2</sub> rooted on cobalt nanoparticles-containing porous carbon polyhedral with dense carbon nanotubes backbone (ZIF-67-C@MoS<sub>2</sub> nanocomposite). The fabrication process for the ZIF-67-C@MoS<sub>2</sub> composite is shown in **Figure 1**. During the preparation process, the ZIF-67 serves not only as the carbon and nitrogen sources for the growth of CNTs, but also as the self-template for the

polyhedron structure. The obtained ZIF-67-C retained a well-defined polyhedral morphology with dense CNTs rooted from both the inner and outer surface of the polyhedron. Then ultrathin MoS<sub>2</sub> covered uniformly on the surface of the ZIF-67-C polyhedron via a facile hydrothermal approach. The *in-situ* grown CNTs and cobalt particles could form an integrated conductive network, which could express the issues of MoS<sub>2</sub> anode including electronic conductivity and volume expansion upon cycling. Furthermore, the robust hollow structure ensures an efficient interfacial contact between active material and electrolyte, beneficial for fast charge-discharge properties.

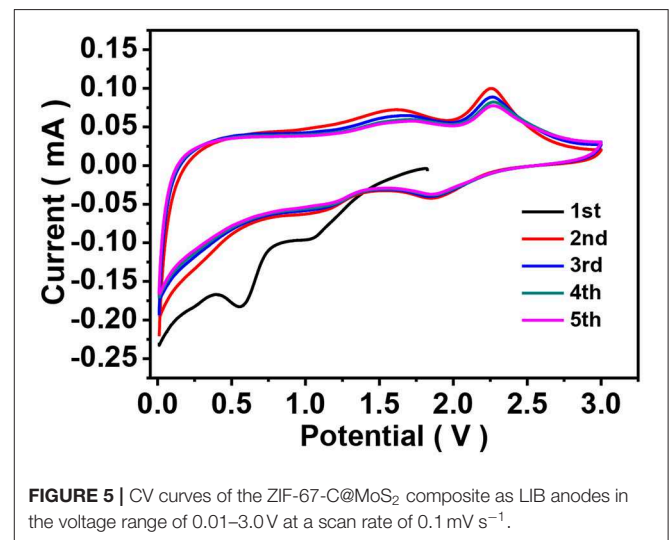
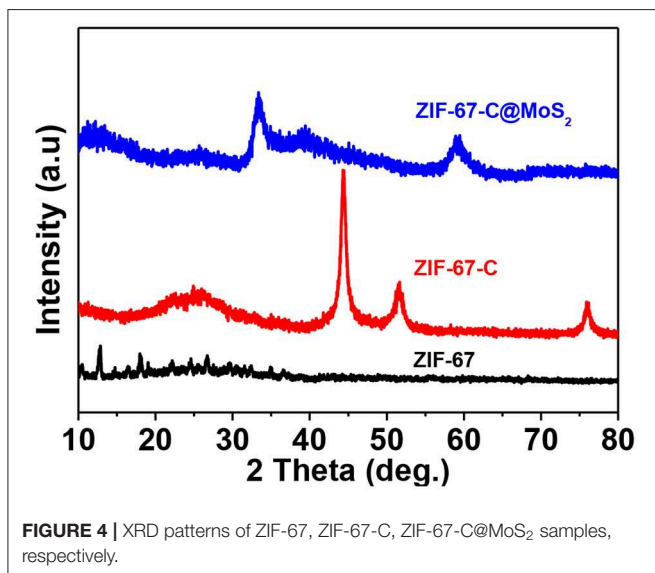
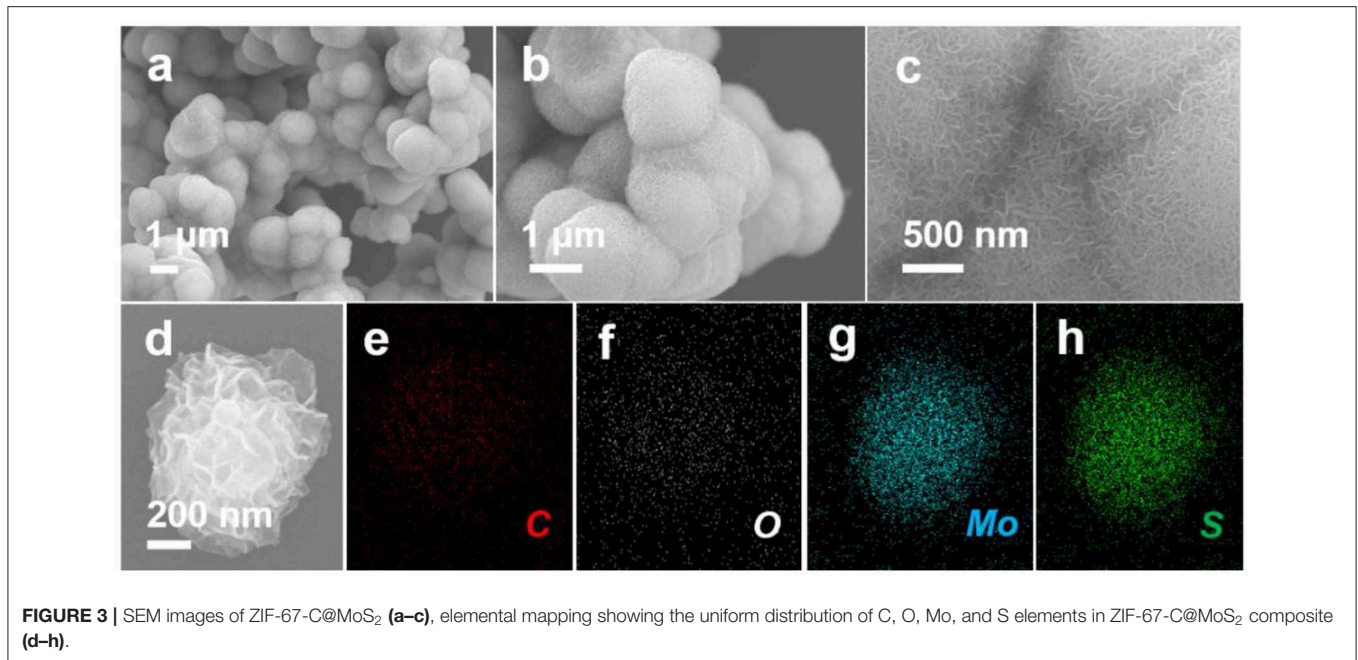
The morphologies and structures of the as prepared ZIF-67, ZIF-67-C, and ZIF-67-C@MoS<sub>2</sub> composites were examined by field emission scanning electron microscopy (FE-SEM). From the SEM images of ZIF-67 (**Figures 2a,b**), it can be seen that ZIF-67 possesses a well-defined polyhedral morphology with a particle size of *ca.* 1–1.5  $\mu\text{m}$ , and each rhombic dodecahedron has a smooth surface, consistent with previous reports in literature (Jiang et al., 2013). As shown in **Figures 2c,d** of the SEM images of ZIF-67-C, the rhombic dodecahedron and size are well-retained after carbonization, and the surface becomes



**FIGURE 2** | SEM images of ZIF-67 (a,b), ZIF-67-C (c,d).



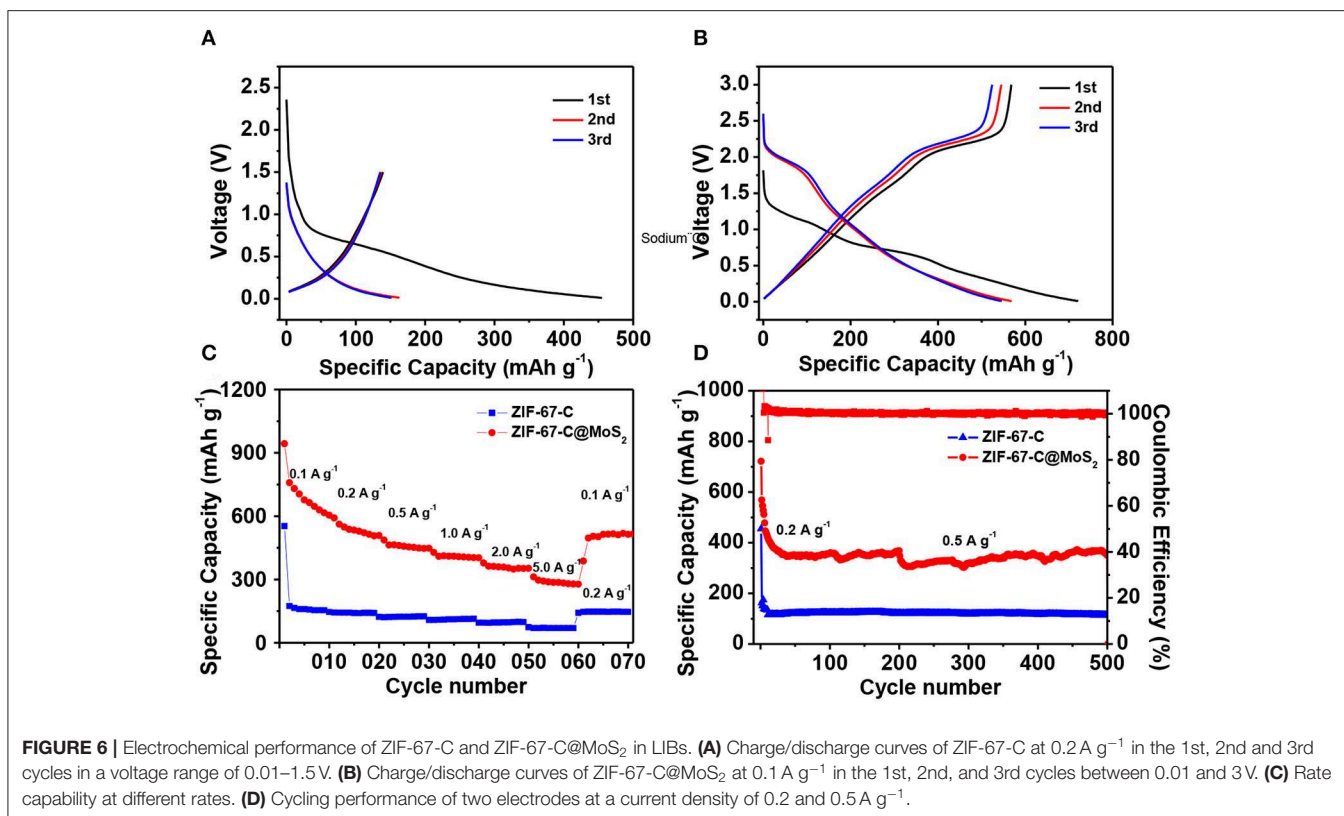
**FIGURE 1** | Schematic illustration for the synthesis of ZIF-67-C@MoS<sub>2</sub> nanocomposite.



rough and many carbon nanotubes covered can be observed. Furthermore, the carbon polyhedron is hollow with dense CNTs rooted on the surface (TEM images are not shown here). During the heat treatment, the H<sub>2</sub> atmosphere plays a critical role in the growth of the carbon nanotubes. Metallic Co nanoparticles are quickly formed in the presence of the H<sub>2</sub> atmosphere, followed by the catalytic growth of carbon nanotubes (Figure 4). In this process, ZIF-67 particles serve not only as the C and N source for growth of carbon nanotubes, but also as the self-template for the framework morphology (Xia et al., 2016). As reported in literature, only solid porous carbon polyhedrons without carbon nanotubes containing cobalt were formatted in argon

or nitrogen atmosphere (Lu et al., 2018). The SEM images of ZIF-67-C@MoS<sub>2</sub> composite are shown in Figures 3a–c. MoS<sub>2</sub> nanosheets were uniformly coated on the surface of ZIF-67-C, and Figure 3c clearly shows the ultrathin MoS<sub>2</sub> flakes on the ZIF-67-C@MoS<sub>2</sub> surface. Elemental mappings of ZIF-67-C@MoS<sub>2</sub> in Figures 3d–h shows the presence and uniform distribution of C, O, Mo, and S elements in the composite, indicating the uniform distribution of MoS<sub>2</sub> nanosheets on the surface of ZIF-67-C.

To explore the phase composition, Figure 4 shows the XRD patterns of the three materials. In the XRD pattern of ZIF-67, the peaks at about 2θ located at 10.3, 14.7, 16.4, 18.9, 21.1, 24.6, 25.6, 26.7, 29.6 corresponds to the (002), (022), (013), (222), (114),



(233), (224), (134), and (044) plane, respectively (Gross et al., 2012; Lin and Chang, 2015), indicating that the prepared ZIF-67 sample is a pure phase. For the XRD pattern of ZIF-67-C, the peak at about 25.2° corresponds to the carbon peak, and the other peaks at about 44.3°, 51.6°, 75.9° can be attributed to metallic Co. The XRD pattern of ZIF-67-C@MoS<sub>2</sub>, the diffraction peaks at 2θ of 33.2, 39.5, and 59.4 corresponded to the (100), (103), and (110) planes of the MoS<sub>2</sub> (Zhang et al., 2018a).

Lithium storage properties of the as-prepared ZIF-67-C and ZIF-67-C@MoS<sub>2</sub> materials were investigated by galvanostatic discharge-charge and CV tests using CR2032 coin cells. To evaluate the electrochemical activity of the ZIF-67-C@MoS<sub>2</sub> composite, CV measurements were carried out at a scan rate of 0.1 mV s<sup>-1</sup>. As shown in **Figure 5**, during the first cathodic scan, the peak located at 1.04 V can be attributed to the Li<sup>+</sup> intercalation into MoS<sub>2</sub> to form Li<sub>x</sub>MoS<sub>2</sub> accompanied by a phase change of MoS<sub>2</sub> from 2H to 1T. And the peak at 0.56 V corresponds to the conversion reaction of Li<sub>x</sub>MoS<sub>2</sub> into metallic Mo particles and Li<sub>2</sub>S (Teng et al., 2016). The broad peak in the range of 0.5–0.8 V can be attributed to the formation of SEI layers. In the subsequent anodic scans, two peaks at 1.60 and 2.25 V are observed. The weak peak at 1.60 V can be assigned to a partial oxidation of Mo to form MoS<sub>2</sub>. The sharp oxidation peak at 2.25 V indicates delithiation of Li<sub>2</sub>S to sulfur. In the following scan, the cathodic peak shifts to 1.87 V, corresponding to the a conversion of S to polysulfide and the formation of Li<sub>2</sub>S. Other peak at 1.20 V can be ascribed to lithiation with metal Mo nanoparticles.

**Figure 6A** shows the discharge/charge curves of the ZIF-67-C sample cycling in the potential range of 0.01–1.5 V at a constant current density of 0.2 A g<sup>-1</sup>. The initial discharge capacity is 454.9 mAh g<sup>-1</sup>, the initial charge capacity is 139.2 mAh g<sup>-1</sup>, corresponding a low Coulombic efficiency (CE) of 30.6%, which can be contributed to the formation of SEI film in the first discharge process. The reversible capacity in the second and third cycle is 163 and 151.5 mAh g<sup>-1</sup>, respectively. The electrode shows one sloping voltage plateau at about 0.31 V (vs. Li<sup>+</sup>/Li) (Jiao et al., 2017), typical charge-discharge curves for carbon material. As shown in **Figure 6B**, the ZIF-67-C@MoS<sub>2</sub> composite delivers the highest specific discharge capacity of 721.2 mAh g<sup>-1</sup> at 0.1 A g<sup>-1</sup>, and a high efficiency of 78.7%. Furthermore, the electrode exhibits a higher reversible capacities of 568.5 and 545.9 mAh g<sup>-1</sup> in the second and third cycles at the same current density.

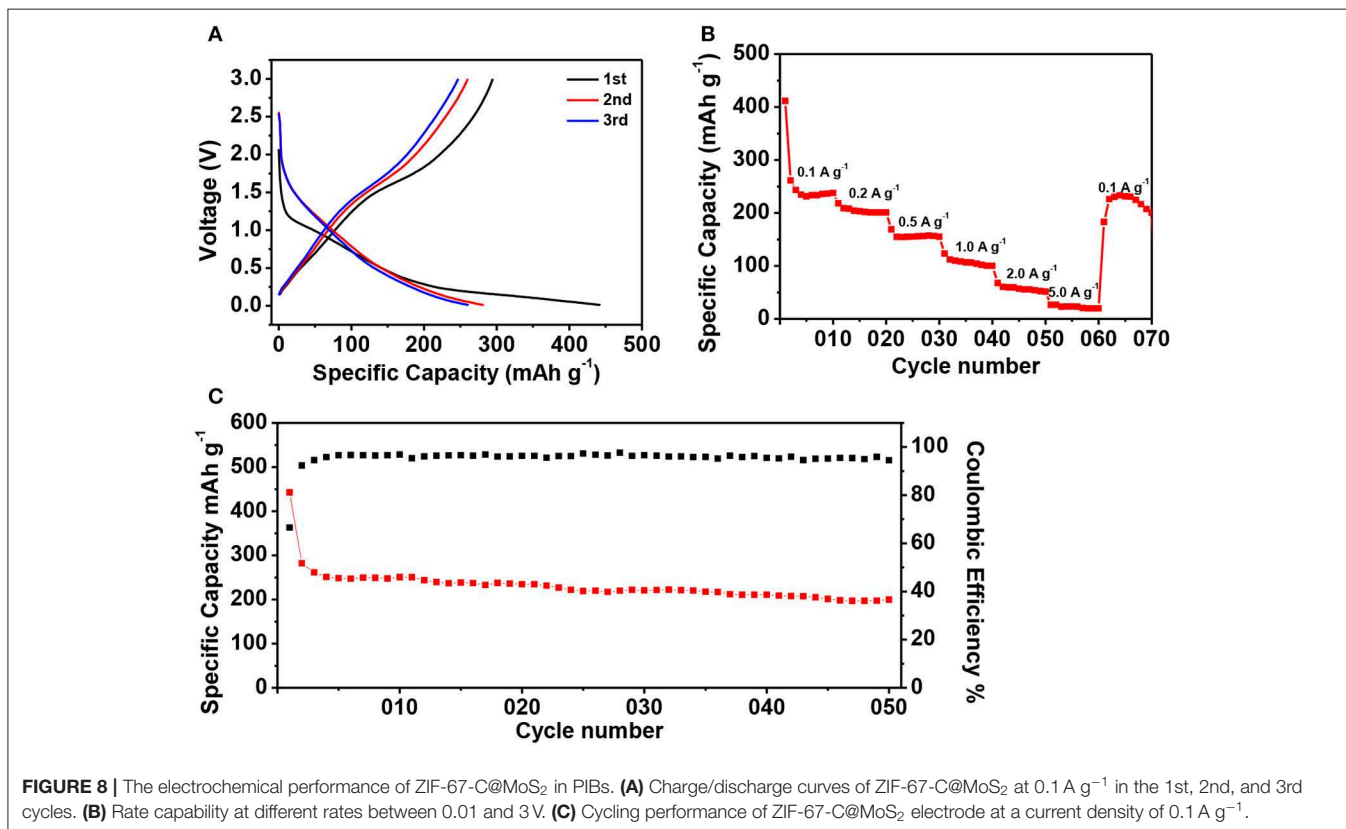
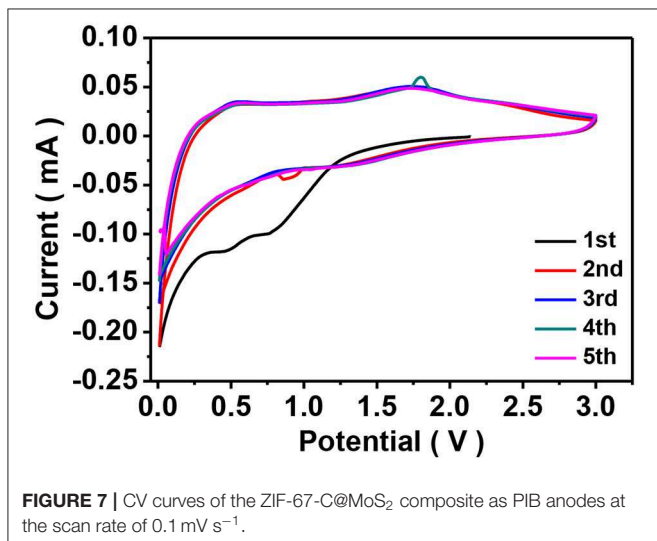
**Figure 6C** compares the rate performance of the two materials at different current rates. It is clear that the reversible capacity of ZIF-67-C@MoS<sub>2</sub> composite is much better than that of ZIF-67-C at various current density. At a low rate of 0.1 A g<sup>-1</sup>, the ZIF-67-C@MoS<sub>2</sub> delivered a discharge capacity of 758.7 mAh g<sup>-1</sup>. Furthermore, the ZIF-67-C@MoS<sub>2</sub> sample could deliver discharge capacities of 561.7, 462.8, 409.3, and 362.8 mAh g<sup>-1</sup> at the current rates of 0.2, 0.5, 1, and 2 A g<sup>-1</sup>, implying the excellent rate capability of the ZIF-67-C@MoS<sub>2</sub> electrode. Even at high rate of 5 A g<sup>-1</sup>, a high discharge capacity of 296.7 mAh g<sup>-1</sup> can still be obtained. When the current density returns to 0.1 A g<sup>-1</sup>, the discharge capacity is 504.7 mAh g<sup>-1</sup>, corresponding to 89.8% retention of the discharge capacity.

The cycling performance of the two materials is also examined. As shown in **Figure 6D**, all electrodes showed very stable capacity retention at a current density of 0.2 and 0.5 A g<sup>-1</sup>. For ZIF-67-C@MoS<sub>2</sub>, the first five cycles are activated at a low current density of 0.1 A g<sup>-1</sup>, then cycling at 0.2 A g<sup>-1</sup> to 200 cycles, and further tested at 0.5 A g<sup>-1</sup> for another 300 cycles. The ZIF-67-C@MoS<sub>2</sub> exhibits a discharge capacity of 350.4 mAh g<sup>-1</sup> after 500 cycles. In contrast, the capacity of ZIF-67-C electrode

falls to lower than 114.2 mAh g<sup>-1</sup> at 0.5 A g<sup>-1</sup> (activating at 0.2 A g<sup>-1</sup> for 10 cycles). The enhanced cycling stability of the ZIF-67-C@MoS<sub>2</sub> samples results from the MoS<sub>2</sub> coating, and synergistic effect of merits of ZIF-67-C, which could efficiently improve electronic conductivity and increased specific capacity.

As mentioned above, the ZIF-67-C@MoS<sub>2</sub> favored fast lithium ions diffusion and improved reaction kinetics with more active sites for the Li<sup>+</sup> insertion/extraction due to the large surface area, effectively increasing the specific capacity of composite. However, it is urgent to find a new generation of energy storage devices due to the limited content of lithium resources in earth. The content of potassium in the earth's crust is 1.5%, which is much higher than that of lithium (0.0017%) (Zhang et al., 2019a). Moreover, the standard reduction potential of K/K<sup>+</sup> (-2.92 V vs. SHE) is close to that of Li/Li<sup>+</sup> (-3.04 V vs. SHE) and lower than that of Na/Na<sup>+</sup> (-2.71 V vs. SHE). Thus, potassium ion batteries (PIBs) have potential application prospect for high energy devices (Chen et al., 2018b; Zhang et al., 2019b).

We furthermore investigated the potassium storage properties of ZIF-67-C@MoS<sub>2</sub> as new anodes. The CV curves of ZIF-67-C@MoS<sub>2</sub> composite were firstly tested in the voltage range of 0.01–3.0 V at a scan rate of 0.1 mV s<sup>-1</sup> (**Figure 7**). During the first cathodic sweep, an obvious reduction peak in the range of 1.2–0.5 V can be clearly seen, which corresponds to the potassiation of MoS<sub>2</sub> and K to form the K<sub>x</sub>MoS<sub>2</sub> (Ren et al., 2017), the irreversible decomposition of the electrolyte and formation of the SEI layers. The peak in 0.01–0.5 V can be attributed to the further reaction of K<sub>x</sub>MoS<sub>2</sub> and K<sup>+</sup> to



form Mo particles and K<sub>2</sub>S. During the subsequent anodic scan, the peak at 1.75 V can be attributed to the deintercalation of K-ion and the oxidation of Mo to MoS<sub>2</sub> (Chen et al., 2018d). **Figure 8** shows the potassium ion storage performance of the ZIF-67-C@MoS<sub>2</sub> electrode with metal potassium as counter electrode. The galvanostatic charge/discharge curves are tested in the potential range of 0.01–3 V at 0.1 A g<sup>-1</sup> in **Figure 8A**. The initial discharge and charge capacity is 442.7 and 294.6 mAh g<sup>-1</sup>, corresponding to an initial CE of 66.5%. The large capacity loss in the first cycle is mainly due to the decomposition of the electrolyte and the formation of the SEI layer. The discharge capacity is 282.0 and 261.2 mAh g<sup>-1</sup> in the second and third cycles, respectively. The electrode shows one voltage plateau at about 0.75 V (vs. K<sup>+</sup>/K). The rate capacity of the ZIF-67-C@MoS<sub>2</sub> electrode is illustrated at various current densities from 0.1, 0.2, 0.5, 1, 2, and 5 A g<sup>-1</sup>, as shown in **Figure 8B**. They deliver high reversible discharge capacities of 261.5, 208.8, 155.1, 112.3, 60.5, and 26.4 mAh g<sup>-1</sup>, respectively. Moreover, when back to 0.1 A g<sup>-1</sup>, the ZIF-67-C@MoS<sub>2</sub> electrode still shows a capacity of 226.3 mAh g<sup>-1</sup>, demonstrating an excellent rate performance. The cycling performance of the ZIF-67-C@MoS<sub>2</sub> is also examined as shown in **Figure 8C**, the initial discharge capacity is 442.7 mAh g<sup>-1</sup>, and delivers a high discharge capacity of 199.8 mAh g<sup>-1</sup> after 50 cycles at a current density of 0.1 A g<sup>-1</sup> with the CE of close to 100%. All in all, the composite exhibits superior rate performance and cycling stability in both lithium-ion and potassium-ion batteries, showing promising potential for large energy storage application.

## CONCLUSIONS

In summary, ultrathin MoS<sub>2</sub> rooted on cobalt nanoparticles-containing porous carbon polyhedral with dense carbon nanotubes has been synthesized by using ZIF-67 as self-template and precursor. During the pyrolysis process, ZIF-67 not only provides the C and N source for growth of N-doped carbon nanotubes arrays catalyzed by the *in-situ* formed metallic Co nanoparticles but also serves as the template for the formation of the hollow polyhedron. The obtained ZIF-67-C@MoS<sub>2</sub> composite exhibits improved specific capacity and

excellent rate capabilities compared to the pristine ZIF-67-C electrode for lithium ion batteries and potassium ion batteries. In particular, the ZIF-67-C@MoS<sub>2</sub> shows the superior lithium storage properties, delivering high reversible capacities of 568.5 mAh g<sup>-1</sup> at 0.2 A g<sup>-1</sup>, superior rate capability (high discharge capacity of 291.9 mAh g<sup>-1</sup> at 5 A g<sup>-1</sup>), and excellent cycling stability for over 500 cycles. The enhanced electrochemical performance of the ZIF-67-C@MoS<sub>2</sub> material could be attributed to the beneficial effects of highly-conductive carbon network with nitrogen doped CNTs and cobalt particle, large specific surface and robust hollow structure. The results here provide a facile strategy for large-scale synthesis of high performance anode materials through low cost perspective, and new insights for designing MOF-derived functional materials with unique structures for energy storage application.

## DATA AVAILABILITY STATEMENT

All datasets generated for this study are included in the article.

## AUTHOR CONTRIBUTIONS

BR, LC, and PN conceived the project and wrote the manuscript. BR carried out the materials synthesis. BR, JL, and HW conducted the materials characterization and the electrochemical evaluation of lithium ion batteries. JL, YG, and LL conducted the tests of potassium ion batteries and the related data processing. All authors discussed the results, read, and commented on the manuscript.

## FUNDING

This work was supported by the National Key R&D Program of China (No. SQ2017YFGH001474), National Natural Science Foundation of China (Nos. 51802111, 51778268), Research Program on Science and Technology from the Education Department of Jilin Province (No. JJKH20190997KJ), Funding of Research Program of Jilin Normal University, JLNU Innovation Program for Graduate Education (No. 201919), and Science and Technology Development Project of Siping City (No. 2017053).

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**Conflict of Interest:** The authors declare that the research was conducted in the absence of any commercial or financial relationships that could be construed as a potential conflict of interest.

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